

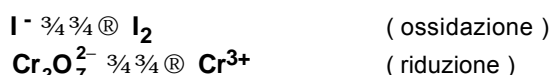
### Bilanciamento delle redox con il metodo delle semireazioni

Bilanciare una reazione significa attribuire ad ogni sostanza presente i coefficienti stechiometrici, in modo che sia possibile la conservazione della massa e la conservazione delle cariche elettriche.

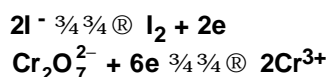
In altre parole il numero di atomi, per ogni specie chimica, presente nei reagenti deve essere uguale al numero di atomi della stessa specie chimica presente nei prodotti di reazione; la carica elettrica complessiva delle sostanze reagenti deve essere uguale alla carica complessiva dei prodotti.

Le procedure per il bilanciamento sono varie; quella che si basa sul metodo delle semireazioni o metodo ionico - elettronico può essere così descritta, utilizzando la reazione in **ambiente acido** già proposta:

1) - Si scrivono separatamente le semireazioni di ossidazione e di riduzione:

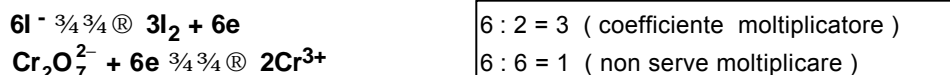


2) - Si bilanciano gli atomi e gli ioni si indicano gli elettroni in movimento:

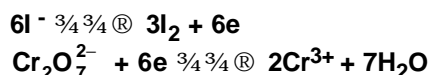


3) - Se il numero di elettroni in gioco nelle due semireazioni non è uguale, si calcola il m.c.m. ( minimo comune multiplo ) dei due valori e lo si divide per il numero di elettroni in ogni semireazione. Il coefficiente ottenuto deve essere moltiplicato per il numero degli elettroni, degli atomi e degli ioni di ciascuna semireazione:

m.c.m. tra 6 e 2 = 6 ; si divide questo valore per il numero degli elettroni nelle due semireazioni:



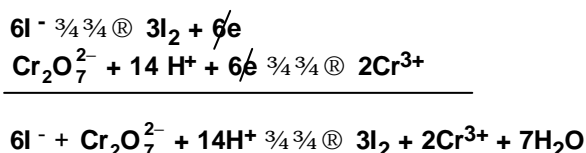
4) - Essendo presente dell'ossigeno, è necessario bilanciarlo con delle molecole di H<sub>2</sub>O:



5) - Si bilancia l'idrogeno dell'acqua con degli ioni H<sup>+</sup> ( ambiente acido ); si controlla il bilanciamento delle cariche:



6) - Si esegue la somma algebrica delle due semireazioni, effettuando le necessarie semplificazioni:



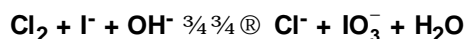
La reazione ionica è così bilanciata.

A questo punto è possibile scrivere la reazione bilanciata anche in forma molecolare:

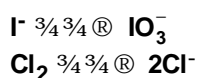


La stessa procedura deve essere utilizzata anche nel caso di reazioni in **ambiente basico**:

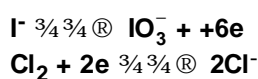
Si consideri la reazione in ambiente basico:



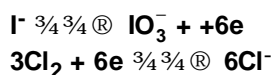
1) - Si scrivono separatamente le semireazioni di ossidazione e di riduzione:



2) - Si bilanciano gli atomi e si indicano gli elettroni in movimento:

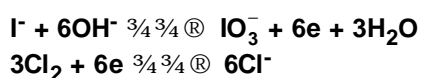


3) - Si calcola il m.c.m. e con tale valore si bilanciano gli elettroni, modificando il numero degli atomi; m.c.m. tra 6 e 2 = 6:



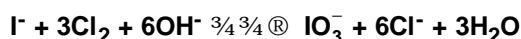
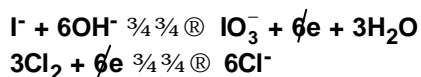
$6 : 6 = 1$ ( non serve moltiplicare ) $6 : 2 = 3$ ( coefficiente moltiplicatore )
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4) - Si bilancia l'ossigeno con molecole di  $\text{H}_2\text{O}$  e l'idrogeno dell'acqua con ioni  $\text{OH}^-$  ( ambiente basico ), controllando il bilanciamento delle cariche:

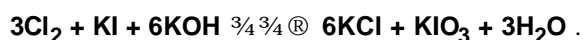


cariche: $-7 \longrightarrow -7$ cariche: $-6 \longrightarrow -6$
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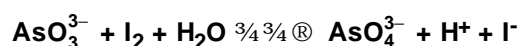
5) - Si somma membro a membro con le opportune semplificazioni:



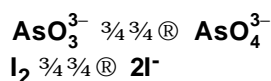
6) - Se richiesto si trasforma la reazione ionica netta in reazione molecolare bilanciata:



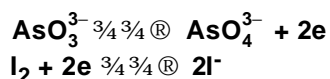
Lo stesso procedimento si utilizza in reazioni che avvengono in **ambiente neutro**:



1) - Si scrivono separatamente le semireazioni di ossidazione e di riduzione:

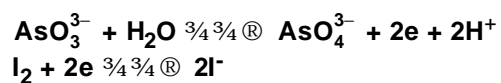


2) - Si bilanciano gli atomi e si indicano gli elettroni in movimento:

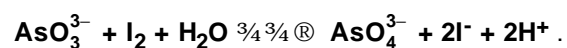
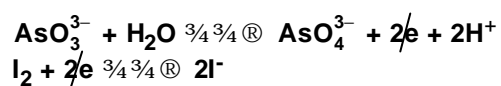


3) - Essendo uguale il numero degli elettroni in movimento non si deve effettuare alcun bilanciamento.

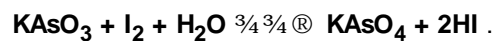
- 4) - Si bilancia nei reagenti l'ossigeno  $\text{H}_2\text{O}$  e nei prodotti l'idrogeno, così aggiunto, con ioni  $\text{H}^+$ , controllando il bilanciamento delle cariche:



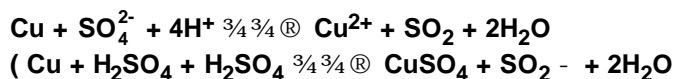
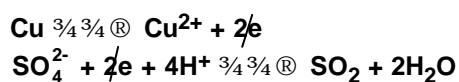
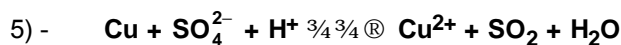
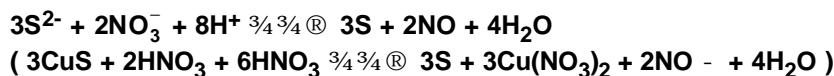
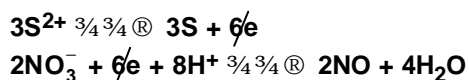
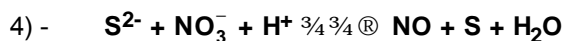
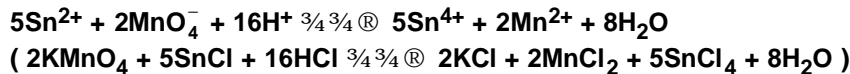
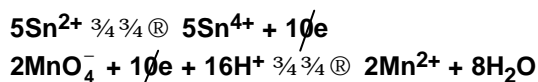
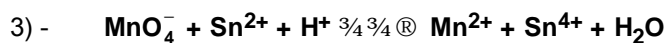
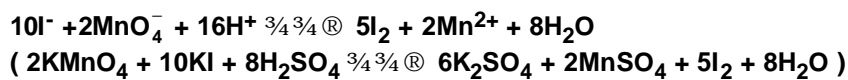
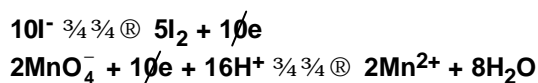
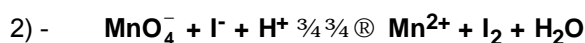
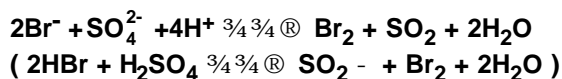
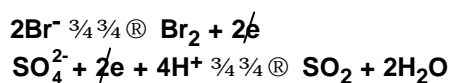
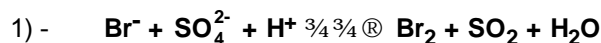
- 5) - Si somma membro a membro con le opportune semplificazioni:

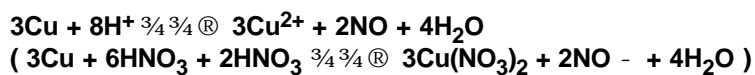
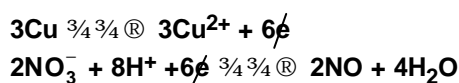
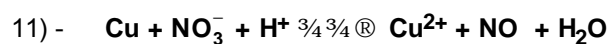
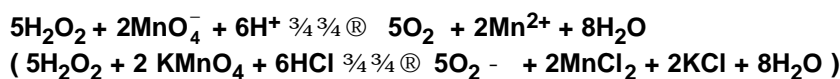
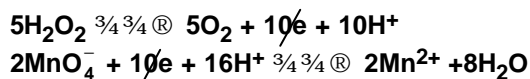
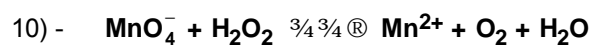
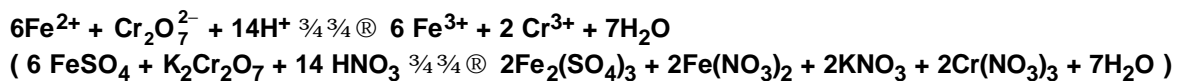
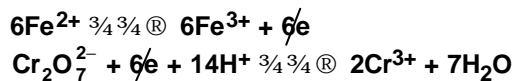
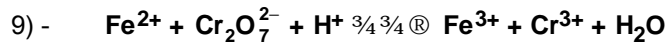
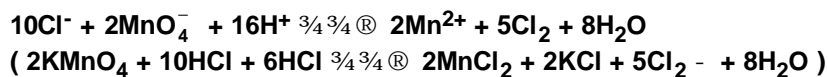
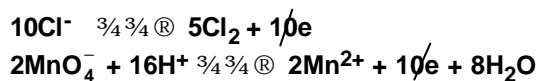
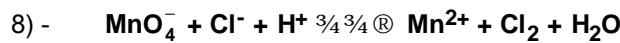
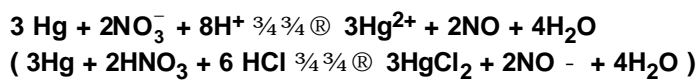
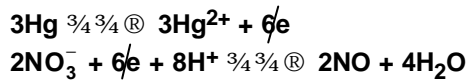
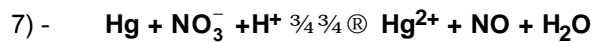
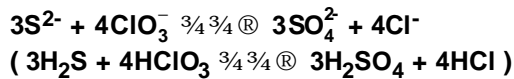
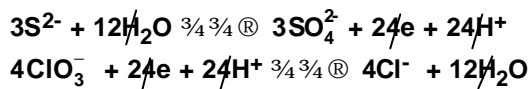
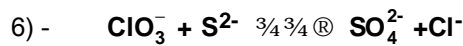


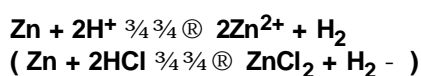
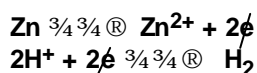
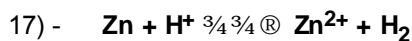
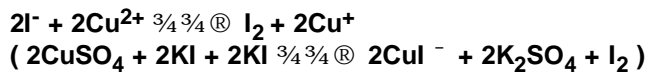
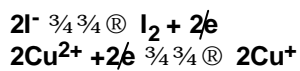
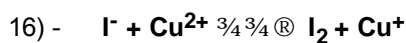
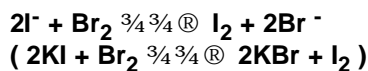
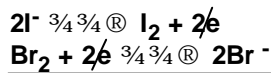
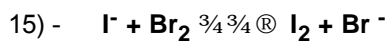
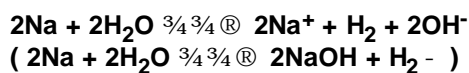
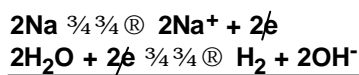
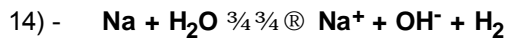
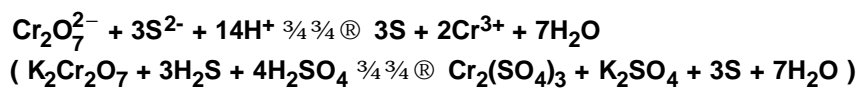
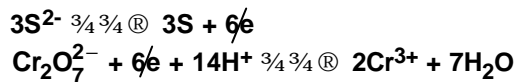
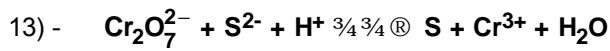
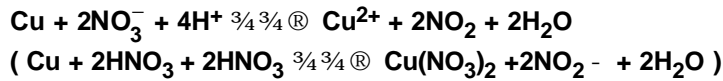
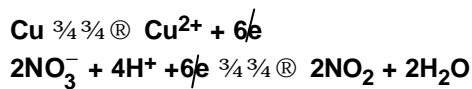
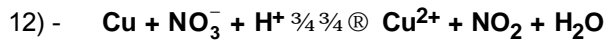
- 6) - Se richiesto si trasforma la reazione ionica netta in reazione molecolare bilanciata:

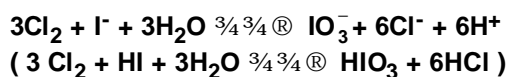
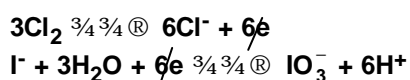
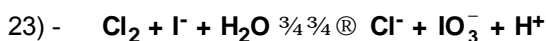
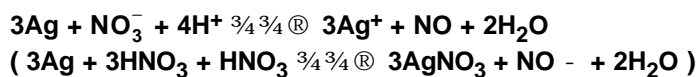
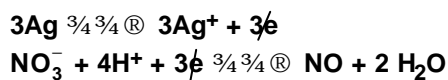
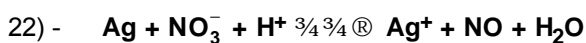
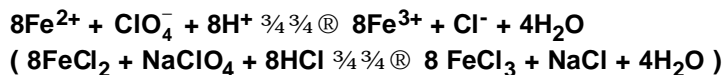
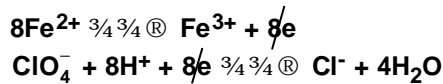
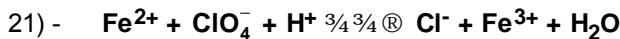
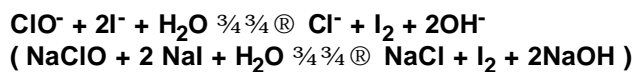
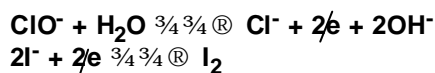
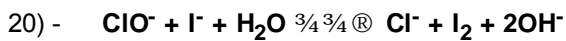
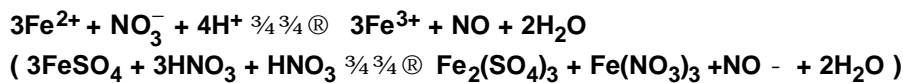
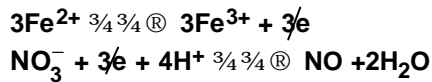
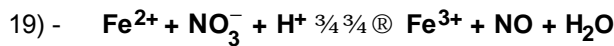
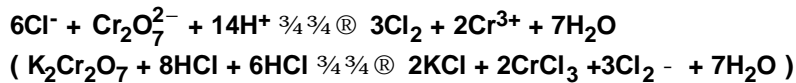
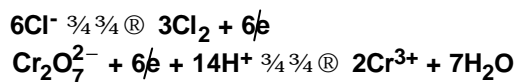
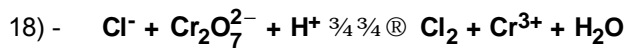


## Esempi di reazioni redox e loro bilanciamento

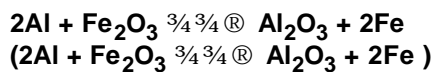
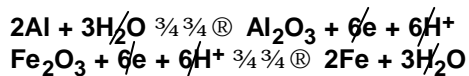
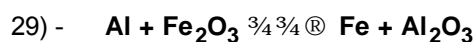
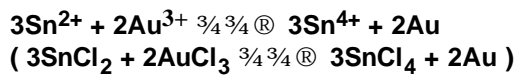
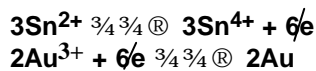
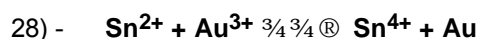
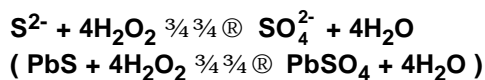
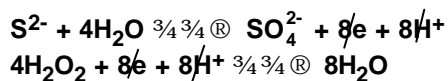
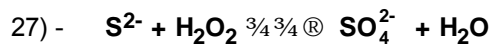
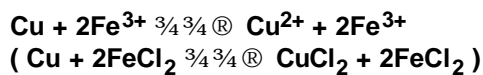
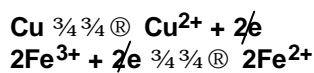
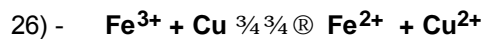
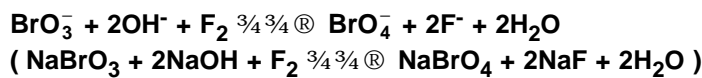
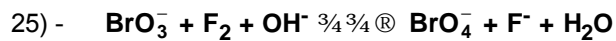
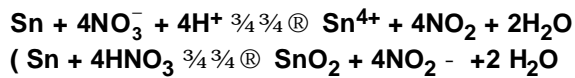
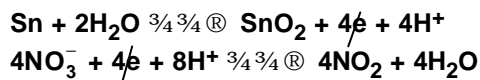
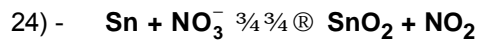


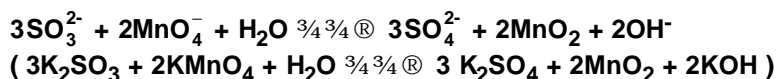
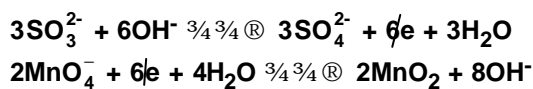
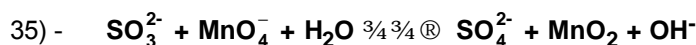
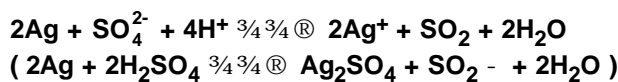
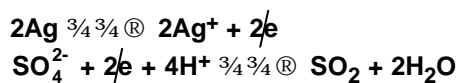
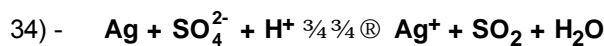
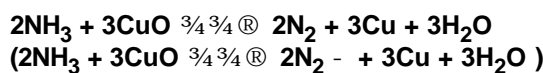
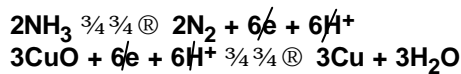
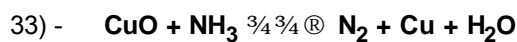
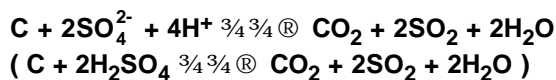
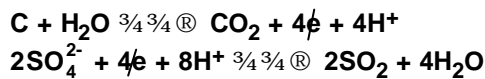
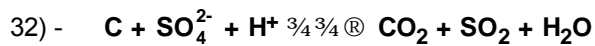
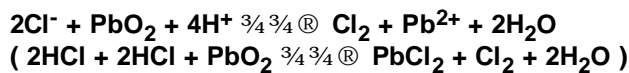
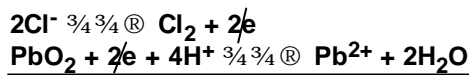
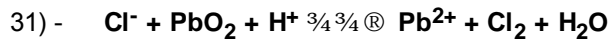
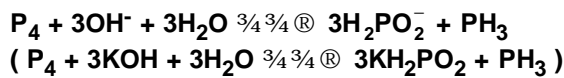
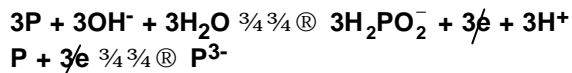
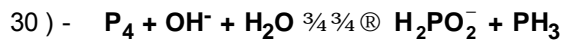


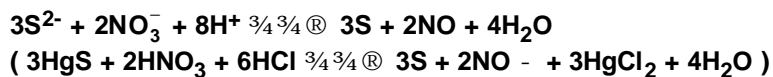
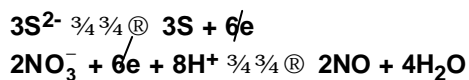
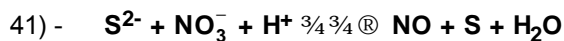
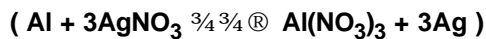
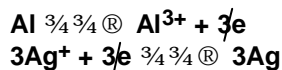
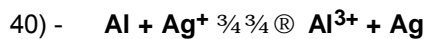
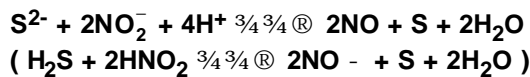
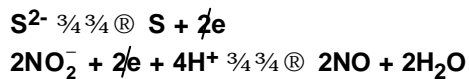
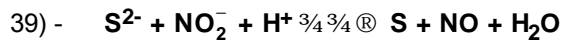
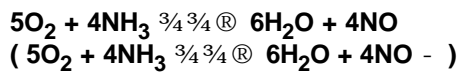
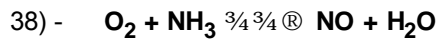
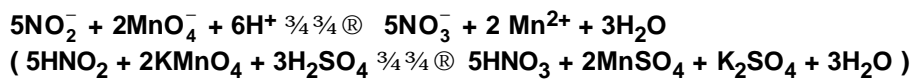
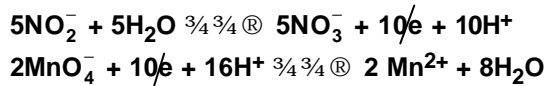
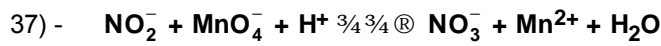
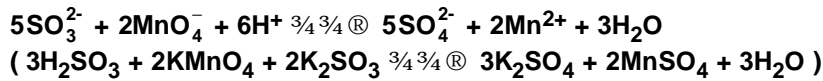
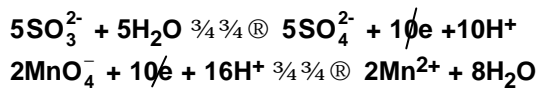
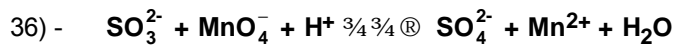


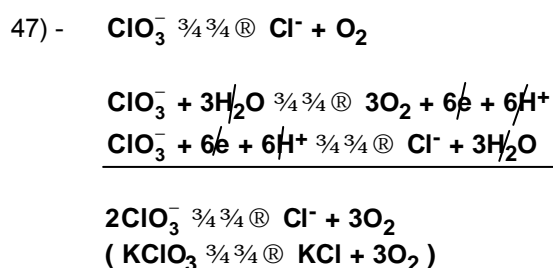
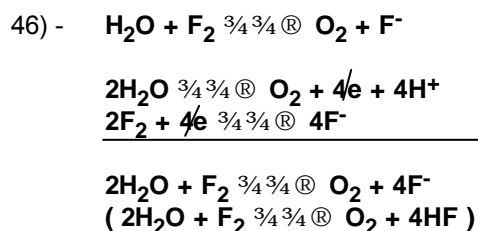
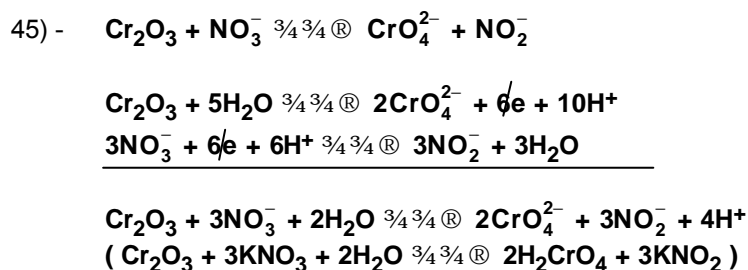
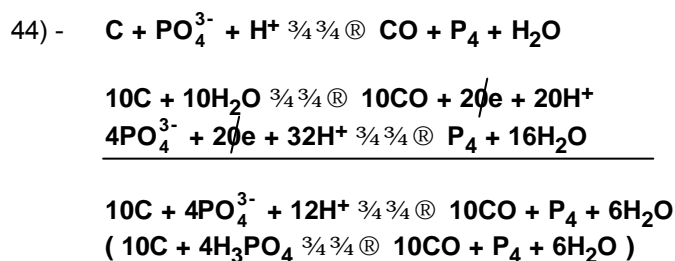
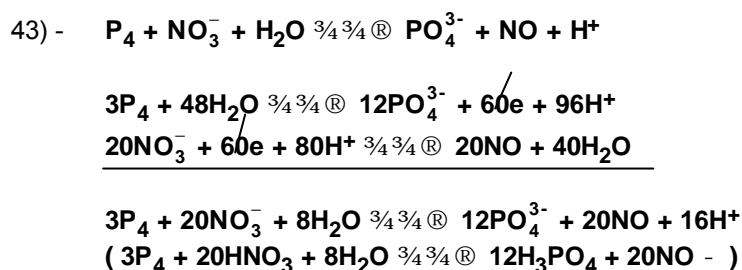
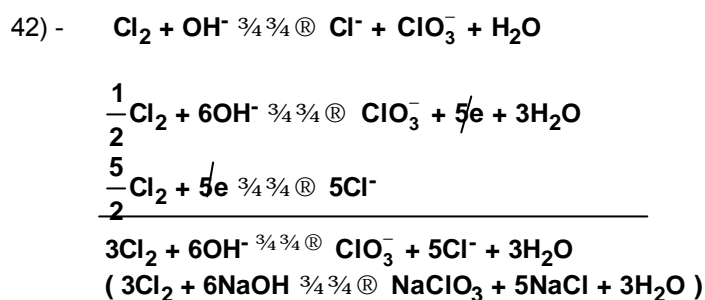


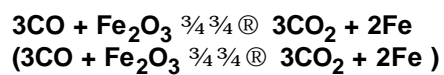
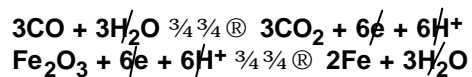
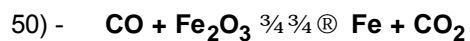
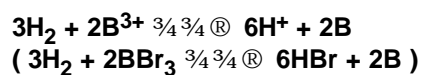
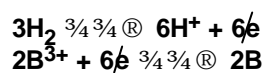
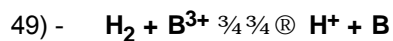
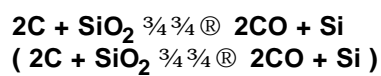
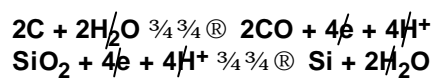
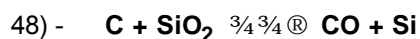












**Nota:** nelle semireazioni ioniche non sono indicati né i bilanciamenti parziali né i calcoli parziali degli elettroni. Il numero di elettroni indicato è quello risultante dal calcolo del m.c.m., come indicato nelle spiegazioni.

Le reazioni molecolari sono, in alcuni casi, solamente teoriche ed esemplificative.

## Riepilogo delle reazioni

- 1) -  $\text{Br}^- + \text{SO}_4^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Br}_2 + \text{SO}_2 + \text{H}_2\text{O}$
- 2) -  $\text{MnO}_4^- + \text{I}^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Mn}^{2+} + \text{I}_2 + \text{H}_2\text{O}$
- 3) -  $\text{MnO}_4^- + \text{Sn}^{2+} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Mn}^{2+} + \text{Sn}^{4+} + \text{H}_2\text{O}$
- 4) -  $\text{S}^{2-} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{NO} + \text{S} + \text{H}_2\text{O}$
- 5) -  $\text{Cu} + \text{SO}_4^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Cu}^{2+} + \text{SO}_2 + \text{H}_2\text{O}$
- 6) -  $\text{ClO}_3^- + \text{S}^{2-} \xrightarrow{\text{3/4 3/4}} \text{SO}_4^{2-} + \text{Cl}^-$
- 7) -  $\text{Hg} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Hg}^{2+} + \text{NO} + \text{H}_2\text{O}$
- 8) -  $\text{MnO}_4^- + \text{Cl}^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Mn}^{2+} + \text{Cl}_2 + \text{H}_2\text{O}$
- 9) -  $\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Fe}^{3+} + \text{Cr}^{3+} + \text{H}_2\text{O}$
- 10) -  $\text{MnO}_4^- + \text{H}_2\text{O}_2 \xrightarrow{\text{3/4 3/4}} \text{Mn}^{2+} + \text{O}_2 + \text{H}_2\text{O}$
- 11) -  $\text{Cu} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Cu}^{2+} + \text{NO} + \text{H}_2\text{O}$
- 12) -  $\text{Cu} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Cu}^{2+} + \text{NO}_2 + \text{H}_2\text{O}$
- 13) -  $\text{Cr}_2\text{O}_7^{2-} + \text{S}^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{S} + \text{Cr}^{3+} + \text{H}_2\text{O}$
- 14) -  $\text{Na} + \text{H}_2\text{O} \xrightarrow{\text{3/4 3/4}} \text{Na}^+ + \text{OH}^- + \text{H}_2$
- 15) -  $\text{I}^- + \text{Br}_2 \xrightarrow{\text{3/4 3/4}} \text{I}_2 + \text{Br}^-$
- 16) -  $\text{I}^- + \text{Cu}^{2+} \xrightarrow{\text{3/4 3/4}} \text{I}_2 + \text{Cu}^+$
- 17) -  $\text{Zn} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Zn}^{2+} + \text{H}_2$
- 18) -  $\text{Cl}^- + \text{Cr}_2\text{O}_7^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Cl}_2 + \text{Cr}^{3+} + \text{H}_2\text{O}$
- 19) -  $\text{Fe}^{2+} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Fe}^{3+} + \text{NO} + \text{H}_2\text{O}$
- 20) -  $\text{ClO}^- + \text{I}^- + \text{H}_2\text{O} \xrightarrow{\text{3/4 3/4}} \text{Cl}^- + \text{I}_2 + 2\text{OH}^-$
- 21) -  $\text{Fe}^{2+} + \text{ClO}_4^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Cl}^- + \text{Fe}^{3+} + \text{H}_2\text{O}$
- 22) -  $\text{Ag} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Ag}^+ + \text{NO} + \text{H}_2\text{O}$
- 23) -  $\text{Cl}_2 + \text{I}^- + \text{H}_2\text{O} \xrightarrow{\text{3/4 3/4}} \text{Cl}^- + \text{IO}_3^- + \text{H}^+$
- 24) -  $\text{Sn} + \text{NO}_3^- \xrightarrow{\text{3/4 3/4}} \text{SnO}_2 + \text{NO}_2$
- 25) -  $\text{BrO}_3^- + \text{F}_2 + \text{OH}^- \xrightarrow{\text{3/4 3/4}} \text{BrO}_4^- + \text{F}^- + \text{H}_2\text{O}$
- 26) -  $\text{Fe}^{3+} + \text{Cu} \xrightarrow{\text{3/4 3/4}} \text{Fe}^{2+} + \text{Cu}^{2+}$
- 27) -  $\text{S}^{2-} + \text{H}_2\text{O}_2 \xrightarrow{\text{3/4 3/4}} \text{SO}_4^{2-} + \text{H}_2\text{O}$
- 28) -  $\text{Sn}^{2+} + \text{Au}^{3+} \xrightarrow{\text{3/4 3/4}} \text{Sn}^{4+} + \text{Au}$
- 29) -  $\text{Al} + \text{Fe}_2\text{O}_3 \xrightarrow{\text{3/4 3/4}} \text{Fe} + \text{Al}_2\text{O}_3$
- 30) -  $\text{P}_4 + \text{OH}^- + \text{H}_2\text{O} \xrightarrow{\text{3/4 3/4}} \text{H}_2\text{PO}_2^- + \text{PH}_3$
- 31) -  $\text{Cl}^- + \text{PbO}_2 + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Pb}^{2+} + \text{Cl}_2 + \text{H}_2\text{O}$
- 32) -  $\text{C} + \text{SO}_4^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$
- 33) -  $\text{CuO} + \text{NH}_3 \xrightarrow{\text{3/4 3/4}} \text{N}_2 + \text{Cu} + \text{H}_2\text{O}$
- 34) -  $\text{Ag} + \text{SO}_4^{2-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{Ag}^+ + \text{SO}_2 + \text{H}_2\text{O}$
- 35) -  $\text{SO}_3^{2-} + \text{MnO}_4^- + \text{H}_2\text{O} \xrightarrow{\text{3/4 3/4}} \text{SO}_4^{2-} + \text{MnO}_2 + \text{OH}^-$
- 36) -  $\text{SO}_3^{2-} + \text{MnO}_4^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{SO}_4^{2-} + \text{Mn}^{2+} + \text{H}_2\text{O}$
- 37) -  $\text{NO}_2^- + \text{MnO}_4^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{NO}_3^- + \text{Mn}^{2+} + \text{H}_2\text{O}$
- 38) -  $\text{O}_2 + \text{NH}_3 \xrightarrow{\text{3/4 3/4}} \text{NO} + \text{H}_2\text{O}$
- 39) -  $\text{S}^{2-} + \text{NO}_2^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{S} + \text{NO} + \text{H}_2\text{O}$
- 40) -  $\text{Al} + \text{Ag}^+ \xrightarrow{\text{3/4 3/4}} \text{Al}^{3+} + \text{Ag}$
- 41) -  $\text{S}^{2-} + \text{NO}_3^- + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{NO} + \text{S} + \text{H}_2\text{O}$
- 42) -  $\text{Cl}_2 + \text{OH}^- \xrightarrow{\text{3/4 3/4}} \text{Cl}^- + \text{ClO}_3^- + \text{H}_2\text{O}$
- 43) -  $\text{P}_4 + \text{NO}_3^- + \text{H}_2\text{O} \xrightarrow{\text{3/4 3/4}} \text{PO}_4^{3-} + \text{NO} + \text{H}^+$
- 44) -  $\text{C} + \text{PO}_4^{3-} + \text{H}^+ \xrightarrow{\text{3/4 3/4}} \text{CO} + \text{P}_4 + \text{H}_2\text{O}$
- 45) -  $\text{Cr}_2\text{O}_3 + \text{NO}_3^- \xrightarrow{\text{3/4 3/4}} \text{CrO}_4^{2-} + \text{NO}_2^-$
- 46) -  $\text{H}_2\text{O} + \text{F}_2 \xrightarrow{\text{3/4 3/4}} \text{O}_2 + \text{F}^-$
- 47) -  $\text{ClO}_3^- \xrightarrow{\text{3/4 3/4}} \text{Cl}^- + \text{O}_2$
- 48) -  $\text{C} + \text{SiO}_2 \xrightarrow{\text{3/4 3/4}} \text{CO} + \text{Si}$
- 49) -  $\text{H}_2 + \text{B}^{3+} \xrightarrow{\text{3/4 3/4}} \text{H}^+ + \text{B}$
- 50) -  $\text{CO} + \text{Fe}_2\text{O}_3 \xrightarrow{\text{3/4 3/4}} \text{Fe} + \text{CO}_2$

## Tavola dei potenziali standard di riduzione ( a 25 °C )

E° ( volts )	reazione elettrodica
+2.87	$F_2(g) + 2e \rightleftharpoons 2F^-(aq)$
+2.02	$S_2O_8^{2-}(aq) + 2e \rightleftharpoons 2SO_4^{2-}(aq)$
+1.77	$H_2O_2(aq) + 2H^+ + 2e \rightleftharpoons 2H_2O$
+1.70	$MnO_4^-(aq) + 4H^+(aq) + 3e \rightleftharpoons MnO_2(s) + 2H_2O$
+1.69	$PbO_2(s) + SO_4^{2-}(aq) + 4H^+ + 2e \rightleftharpoons PbSO_4(s) + 2H_2O$
+1.49	$MnO_4^-(aq) + 8H^+(aq) + 5e \rightleftharpoons Mn^{2+}(aq) + 4H_2O$
+1.45	$PbO_2(s) + 4H^+(aq) + 2e \rightleftharpoons Pb^{2+}(aq) + 2H_2O$
+1.44	$BrO_3^-(aq) + 6H^+(aq) + 6e \rightleftharpoons Br^-(aq) + 3H_2O$
+1.42	$Au^{3+}(aq) + 3e \rightleftharpoons Au(s)$
+1.36	$Cl_2(g) + 2e \rightleftharpoons 2Cl^-(aq)$
+1.33	$Cr_2O_7^{2-}(aq) + 14H^+(aq) + 6e \rightleftharpoons 2Cr^{3+}(aq) + 7H_2O$
+1.23	$O_2(g) + 4H^+ + 4e \rightleftharpoons 2H_2O$
+1.20	$Pt^{2+}(aq) + 2e \rightleftharpoons Pt(s)$
+1.19	$IO_3^-(aq) + 6H^+(aq) + 5e \rightleftharpoons \frac{1}{2}I_2(s) + 9H_2O$
+1.07	$Br_2(aq) + 2e \rightleftharpoons 2Br^-(aq)$
+0.96	$NO_3^-(aq) + 4H^+(aq) + 3e \rightleftharpoons NO(g) + 2H_2O$
+0.90	$Hg^{2+}(aq) + 2e \rightleftharpoons Hg(s)$
+0.87	$NO_3^-(aq) + 4H^+(aq) + 2e \rightleftharpoons 2NO_2(g) + 2H_2O$
+0.80	$Ag^+(aq) + e \rightleftharpoons Ag(s)$
+0.77	$Fe^{3+}(aq) + e \rightleftharpoons Fe^{2+}(aq)$
+0.69	$O_2(g) + 2H^+(aq) + 2e \rightleftharpoons H_2O_2(aq)$
+0.53	$I_2(s) + 2e \rightleftharpoons 2I^-(aq)$
+0.45	$SO_2(aq) + 4H^+(aq) + 4e \rightleftharpoons S(s) + 2H_2O$
+0.401	$O_2(g) + 2H_2O + 4e \rightleftharpoons 4OH^-(aq)$
+0.34	$Cu^{2+}(aq) + 2e \rightleftharpoons Cu(s)$
+0.17	$SO_4^{2-}(aq) + 4H^+(aq) + 2e \rightleftharpoons H_2SO_3(aq) + H_2O$
+0.16	$Cu^{2+}(aq) + e \rightleftharpoons Cu^+(aq)$
+0.15	$Sn^{4+}(aq) + 2e \rightleftharpoons Sn^{2+}(aq)$
+0.14	$S(s) + 2H^+(aq) + 2e \rightleftharpoons H_2S(g)$
<b>+0.00</b>	<b><math>2H^+(aq) + 2e \rightleftharpoons H_2(g)</math></b>
-0.13	$Pb^{2+}(aq) + 2e \rightleftharpoons Pb(s)$
-0.14	$Sn^{2+}(aq) + 2e \rightleftharpoons Sn(s)$
-0.25	$Ni^{2+}(aq) + 2e \rightleftharpoons Ni(s)$
-0.28	$Co^{2+}(aq) + 2e \rightleftharpoons Co(s)$
-0.44	$Fe^{2+}(s) + 2e \rightleftharpoons Fe(s)$
-0.74	$Cr^{3+}(aq) + 3e \rightleftharpoons Cr(s)$
-0.76	$Zn^{2+}(aq) + 2e \rightleftharpoons Zn(s)$
-0.83	$2H_2O + 2e \rightleftharpoons H_2(g) + 2OH^-(aq)$
-1.19	$Mn^{2+}(aq) + 2e \rightleftharpoons Mn(s)$
-1.67	$Al^{3+}(aq) + 3e \rightleftharpoons Al(s)$
-2.37	$Mg^{2+}(aq) + 2e \rightleftharpoons Mg(s)$
-2.71	$Na^+(aq) + e \rightleftharpoons Na(s)$
-2.76	$Ca^{2+}(aq) + 2e \rightleftharpoons Ca(s)$
-2.89	$Sr^{2+}(aq) + 2e \rightleftharpoons Sr(s)$
-2.90	$Ba^{2+}(aq) + 2e \rightleftharpoons Ba(s)$
-2.92	$K^+(aq) + e \rightleftharpoons K(s)$
-3.04	$Li^+(aq) + e \rightleftharpoons Li(s)$